

Self Quiz

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2. No.

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2. The gaining, losing, or sharing of neutrons.
3. The gaining, losing, or sharing of electrons.
4. The gaining, losing, or sharing of isotopes.
5. The gaining, losing, or sharing of ions.

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2. A negative ion.
3. A neutral ion.
4. An isotope.
5. A compound.

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How many electrons does sodium lose to chlorine to form an ionic bond?

1. 1

2. 2

3. 3

4. 4

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What is the difference between a nonpolar covalent bond and a polar covalent bond?

1. A polar covalent bond results when there is unequal sharing of electrons in a molecule, whereas electrons are shared equally in a nonpolar covalent bond.
2. A polar covalent bond has two equal sides and a nonpolar covalent bond has two different sides.
3. A nonpolar covalent bond is positively charged and a polar covalent bond is negatively charged.
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A covalent chemical bond is one in which

- A) electrons are removed from one atom and transferred to another atom so that the two atoms become oppositely charged.**
- B) protons or neutrons are shared by two atoms**
- C) electrons of one atom are transferred to another atom.**
- D) electrons are shared by two atoms**

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A covalent bond is likely to be polar when

- A) the electron being shared between two atoms is much more attracted to one of the atoms than the other atom.**
- B) the two atoms sharing electrons are different elements.**
- C) the two atoms sharing electrons are of the same element.**
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The ionic bond of sodium chloride is formed when

- A) sodium and chlorine both lose electrons.**
- B) chlorine gains a proton from sodium.**
- C) sodium gains an electron from chlorine.**
- D) sodium and chlorine share an electron pair.**
- E) chlorine gains an electron from sodium.**

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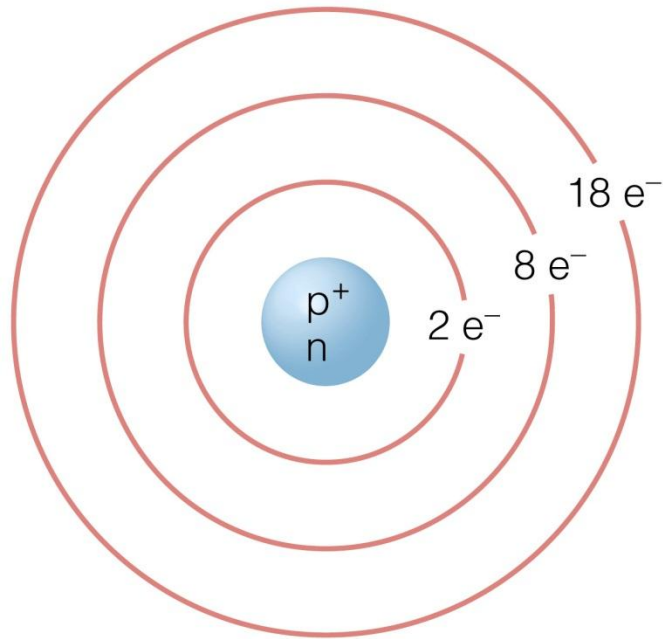
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NOTE: Details of electron orbitals not expected

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